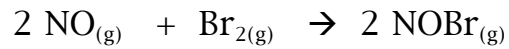


Rate Law Experiment



$$\text{RATE} = k [\text{NO}]^x [\text{Br}_2]^y$$

The reaction was carried out 5 times using different initial reactant concentrations. The initial rate was measured for each experiment.

Experiment	$[\text{NO}]_0$	$[\text{Br}_2]_0$	initial rate $\{-\Delta[\text{NO}]/\Delta t\}$
1	.10	.10	12 M/sec
2	.10	.20	24 M/sec
3	.10	.30	36 M/sec
4	.20	.10	48 M/sec
5	.30	.10	108 M/sec

Determine x, y, and k from data.